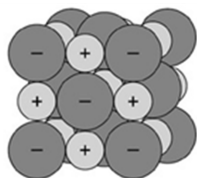
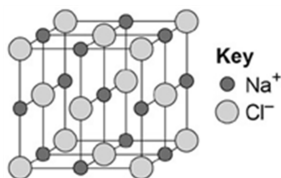


Ionic Compounds

25 November 2019 10:00

An ionic compound is a giant structure of ions. Ionic compounds are held together by strong electrostatic forces of attraction between oppositely charged ions. These forces act in all directions in the lattice and this is called ionic bonding.

The structure of sodium chloride can be represented in the following forms:



Left dia not to scale, & no gaps
Right dia - shows relative sizes, but you can't see inside.

Both go on & on.

All ionic compounds have similar properties

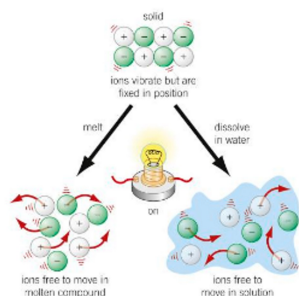


Figure 3 Ionic compounds do not conduct electricity in the solid state but do when molten or when dissolved in water

The ions are "locked" in place so doesn't conduct. But...

- melt it & ions are free to move
- dissolve - ions are free to move

Empirical Formula

Look at the diagram to work out what ions are present

Look at the periodic table to work out what ions are present

Balance them up to make a molecule